

84.5

$$m(\text{CuSO}_4 \cdot 5\text{H}_2\text{O}) = 250 \cdot 0.001$$

$$m(\text{CuSO}_4) = 250 \cdot 0.001$$

$$x = 0.001 \cdot 250 = 0.25 \text{ g}$$

$$m(\text{CuSO}_4) = \frac{250}{100} = 2.5 \text{ g}$$

$$m(\text{CuSO}_4 \cdot 5\text{H}_2\text{O}) = 2.5 \cdot 250 = 625 \text{ g}$$

... 361 + ...



$$m(\text{Cu}_2\text{O}) = 104 \text{ g}$$

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$$x = 54$$

(155)

... ..

③

PH₃



④

(4)



(30)

Окисление x_1 - окислитель
 x_2 - восстановитель
 x_3 - восстановитель
 x_4 - окислитель

